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Stoichiometry. SECTION 1. SHORT ANSWER Answer the following questions in the space provided. 1. _____ The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products. (c) number of atoms of each element in each compound in a reaction.

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*Chapter 9 Section 9.1: Team Learning
Worksheet*

11—1 Review and Reinforcement

Stoichiometry ... stoichiometry problem.
A balanced equation verifies the law of If
the statement is true, write "true. " If it
is false, change the underlined word or
words to make the statement true. Write

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your answer on the line provided s. 10.
12. The term stoichiometry is derived from the words

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11-1 - 11-3 . CH 11 Gas Stoichiometry 1 .
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SECTION 9-3 PROBLEMS Write the

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answer on the line to the left. Show all your work in the space provided.

1. 88%
If the actual yield of a reaction is 22 g and the theoretical yield is 25 g, calculate the percent yield.

2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the following equation:
$$N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$$

 N_2 ; 2.0 mol a.

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Chapter 9 focuses on reaction stoichiometry: using a balanced chemical equation to calculate the number of grams, moles, or particles of reactants/products involved in a chemical reaction. Students had an introduction to composition

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stoichiometry in Chapter 3 and will now move on to some more difficult problems.

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Stoichiometry Review Answers 1. a.

Na₃PO₄ b. Ca(NO₃)₂ ... Use the

following balanced equation. Identify the limiting reactant when 1.150 grams of

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HgO react with 12.46 grams of Cl₂.
Convert to moles to get moles available,
then calculate moles required:

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1. An individual coefficient does not tell us anything. ... coefficient when doing stoichiometry calculations. For example, when given any mass of oxygen ... 1.

The answer is “e” (each would produce the same amount of product). Many students try to answer this question without performing calculations ...

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calculator! Round molar masses to the "ones" place. Chapter 9: Stoichiometry - GO.HRW.COM

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and Reactions Teacher Notes and
Answers Chapter 8 SECTION 1 SHORT

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ANSWER 1. a. d b. a c. b d. f e. e f. c 2.

8,4,9 3. a. 12 atoms ... SECTION 3

SHORT ANSWER 1. Choose from Cu, Ag, Au, Pt, Sb, Bi, and Hg. 2. Fe forms an oxide in nature, and Ag does not, because it is much less active. 3. a. F₂ b. K c. H 4. a.

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SECTION 1 SHORT ANSWER Answer the following questions in the space

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provided. 1. The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products.

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SECTION 1 SHORT ANSWER Answer the following questions in the space provided.

1. b The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products.

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