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Two equilibria must be considered: CuO(s) + H2(g) - H2O(l) + Cu at equilibrium $- nOD2 - Cu(s) + \frac{1}{2}O2(g) - CuO1$ at equilibrium - nOD2 - 2 In the final state, only O2 is present. Therefore = nOD2. In an excess of O2, all the 1 copper is oxidized. Therefore - = 0 or = nOD2.

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