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Quantities Goals : To

gain an understanding

of : 1. Problem solving

in chemistry. 2. The

use of dimensional

analysis to solve

problems. 3. The

concept of the mole. 4.

The relationship

between masses of

substances and moles

of substances. 5. The

relationship between

moles and the volumes

and densities of gases.

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6.

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Notes. By the  
attraction between the  
positive nucleus of one  
atom and the negative  
electrons of another  
atom.

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Chapter 7 7-1 symbol?  
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an abbreviation for the name of the element

Chemical formula? a type of notation made with numbers and

chem. Symbols To indicate the composition of a

compound To indicate the number of atoms in one molecule of an

element molecule? a single atom, a group of 2+ atoms of the same element, or a group of atoms of different elements that have

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combined to form a  
compound Monoatomic  
molecules? molecules  
with one atom

Diatomic molecules?  
two atom molecules  
Hydrogen, oxygen ...

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...

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5 April 11, 2018 review

parts of a formula

$\text{CoSO}_4 \text{ Al}_2(\text{SO}_4)_3$  this 2

means 2 Aluminum

atoms in each unit this

4 means 4 oxygen this

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3 means that the polyatomic ion sulfate is in the unit 3 times therefore, 3 sulfur and 12 oxygen!

cation/anion review

CoSO<sub>4</sub> The rules.....

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Question#7) How can a non-spontaneous reaction be carried out in an electrolytic cell?

Discuss in detail?

Answer: Electrolytic

cells: The type of

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electrochemical cell in which a non-spontaneous (energy is used) chemical reaction takes place when electric current is passed through the solution, is called an electrolytic cell.

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Equilibrium In a chemical reaction chemical equilibrium is defined as the state at which there is no further change in concentration of reactants and products. For example, At equilibrium the rate of forward reaction is equal to the rate of backward reaction.

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Chapter 7 Equilibrium.

Introduction In a reversible reaction, a stage is reached when the concentrations of reactants and products stay unchanged. This state is called as the state of equilibrium. At equilibrium, the concentrations of reactants and products may or may not be equal but must remain unchanged with time.

**Plus One Chemistry**

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## **Notes Chapter 7**

### **Equilibrium - A Plus Topper**

252 Chapter 7 Energy and Chemical Reactions 1 Although chemists recognize that matter can be converted into energy and energy into matter, this matter-energy conversion is small enough to be disregarded.

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## **Introduction to Chemistry: Energy and ...**

This video contain the explanation of some chemical reaction, differentiate between physical changes and chemical reactions, classify the chemical reactions-...

**Chemical Reaction |  
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□□□□□□□□ □□□□□□□□ |

**Fahad Sir**

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does your chapter outline compare to the chapter summary and key terms, key skills, and key equations at the end of each chapter?

## **General Chemistry I - CHM2045**

A homogeneous mixture of different chemical substances which has uniform chemical composition through out and shows uniform physical

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properties is called solution. For example dissolve a small amount of copper sulphate in water the water will become blue.

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Chemistry 108 Lecture Notes Chapter 7:

Solutions 14. Parts per thousand, parts per million, and parts per billion. • These units

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are mostly used for very dilute solutions •  
If 30 mg of sucrose is dissolved in a swimming pool, the sucrose concentration is 0.001 ppm and 1 ppb.

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can give a quick glance  
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