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WebEx Chemistry

Lesson 6.1-6.2

Avogadro /u0026 the

Mole April 6

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Sustainability Lecture

- Connecting the

Drops: Water Scarcity

and the Economy

David Rusenko - How

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To Find Product
Market Fit

~~Introduction to Moles~~

~~Mole Conversions~~

~~Made Easy: How to~~

~~Convert Between~~

~~Grams and Moles~~

~~How to Make Borax~~

~~Crystal Ornaments-~~

~~with yoyomax12~~

Growing Cannabis

With Love With The

Plant Centric Grower

Lovin In Her Eyes

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Avogadro's Number,
The Mole, Grams,
Atoms, Molar Mass
Calculations -

Introduction The Map
of Mathematics

~~Chemistry - 2nd sec
part 1 unit 4 -~~

~~ALKALI metals - atomic
radius - atomic
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~~photoelectric ph~~

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Tradecraft | WIRED~~

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Using Avogadro ' s
Number | How to
Pass Chemistry
Interconverting
Masses, Moles and

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Quantities of Particles

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CHAPTER 1 BASIC

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Unit 7 Review. Click
here to review the
key concepts of this
unit.

Unit 7: Chemical
Quantities - Home
Chemical Quantities
Here you will learn

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Quantities, molar mass, and conversions. What is a Mole? The dictionary defines a mole as: "the SI unit of amount of substance, equal to the quantity containing as many elementary units as there are atoms in 0.012 kg of carbon-12".

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Quantities

Chapter 10: Chemical
Quantities - Weebly

Use the chemical equation to determine how many moles of water are present.

Total the mass for all elements that make up one mole of water. Determine how many grams are present in 2 moles of water.

Convert 36.04 g of

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Quantities
water to kilograms
using the conversion
method.: e y . . The
coefficient in front of
H₂O is 2, so 2 moles
of water are present.

Chemical Quantities - AP Biology

1. Use the chemical
formula to determine
the number of each
type of atom present
in the compound.

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2. Multiply the atomic weight (from the periodic table) of each element by the number of atoms of that element present in the compound.

3. Add it all together and put units of grams/mole after the number.

Chemical Quantities -
Exploring Chemistry

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Mole is a unit of measurement used in chemistry that shows the amount of a chemical substance. To find the molar mass of a compound you will use the chemical formula to determine the number of each type of atom present in the compound, then you multiply the atomic

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weight of each element by the number of atoms of that element present in the compound.

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Chemistry
Chemical Quantities.
Chemical Reactions >
Oxidation-Reduction
Reactions > Solutions.
Acids, Bases, and
Salts > Behavior of

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software that inspires.
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Unit 6 - Quantities in
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Chapter 10: Chemical
Quantities Chapter

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11: Chemical

Reactions Chapter 14:

Behavior of Gases

Chapter 16: Solutions

Chapter 19: Acids,

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20: Oxidation-

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Chemistry Unit 3 -

Chemical quantities

In this unit, students learn about how mass is conserved in chemical reactions and how we can measure amounts of substances. This is looked at in more detail and students will learn about calculating relative formula masses and

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Year 10 Chemistry

unit 3 chemical

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substance. To find the molar mass of a compound you will use the chemical

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formula to determine the number of each type of atom present in the compound, then you multiply the atomic weight of each element by the number of atoms of that element present in the compound.

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Chemistry Page 6/22

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SCH3UI Unit 3

Chemical Quantities -
Mr. Arthur's Science
Page

Chemical Quantities
and Reactions - MS.
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Chapter 10: Chemical
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Mole is a unit of
measurement used in
chemistry that shows
the amount of a

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Quantities substance.

To find the molar mass of a compound you will use the chemical formula to determine the number
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Stoichiometry-
Greek for

"measuring elements",
stoichiometry is the
calculation of
quantities in chemical
reactions is based on
balanced equations.

Significance of a
Balanced Equation-
When an equation is
balanced, the

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Coefficient numbers
represent numbers of
moles or particles
(atoms, molecules or
formula units).

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