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Chapter 7 Review Chemical Formulas

CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Label each of the following statements as True or False: True a. If the formula mass of one molecule is x amu, the molar mass is x g/mol. False b. Samples of equal numbers of moles of two different chemicals

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7 Chemical Formulas and Chemical Compounds

CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Write formulas for the following compounds: CuCO₃ a. copper(II) carbonate Na₂SO₃ b. sodium sulfite (NH₄)₃PO₄ c. ammonium phosphate SnS₂ d. tin(IV) sulfide HNO₂ e. nitrous acid 2.

7 Chemical Formulas and Chemical Compounds

Modern Chemistry 57 Chemical Bonding CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Label each of the following statements as True or False: _____ a. If the formula mass of one molecule is xu , the molar mass is x g/mol. _____ b.

CHAPTER 7 REVIEW Chemical Formulas and Chemical

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Compounds

CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. c In a Stock system name such as iron(III) sulfate, the Roman numeral tells us (a) how many atoms of Fe are in one formula unit. (b) how many sulfate ions can be attached to the iron atom. (c) the charge on ...

7 Chemical Formulas and Chemical Compounds

Chapter 7: Chemical Formulas and Chemical Compounds Section 2: Oxidation Numbers Oxidation numbers An oxidation number, also called oxidation state, indicates the general distribution of electrons among the bonded atoms in a molecular compound or polyatomic ion. Unlike ionic charges, oxidation numbers do not have an exact physical meaning.

Chapter 7: Chemical Formulas and Chemical Compounds

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Chapter 7 - Chemical Formulas & Chemical Compounds Chapter 7 is the final chapter of the first semester. It serves as a summary of all course content discussed so far, as the focus of this chapter is on process skills rather than new concepts or theory.

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CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. Assign the oxidation number to the specified element in each of the following examples: a. S in H_2SO_3 b. S in MgSO_4 c. S in K_2S d. Cu in Cu_2S e. Cr in Na_2CrO_4 f. N in HNO_3 g. C in (HCO_3) h. N in (NH_4) 2. a.

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chemical formulas. In this chapter, you will be introduced to some of the rules used to identify simple chemical compounds.

Significance of a Chemical Formula Recall that a chemical formula indicates the relative number of atoms of each kind in a chemical compound. For a molecular compound, the

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Chapter 6 "Chemical Bonding" Chapter 7 "Chemical Formulas and Chemical Compounds" Chapter 8 "Writing and Balancing Equations" Chapter 9 "Stoichiometry" Chapter 11 "Gas Laws" Chapter 14 "Lewis Dot Structures, Acids & Bases" VSEPR Chapters 12 & 15 "Molarity, pH & Titrations" Chapter 18 "Chemical Equilibrium"

Chapter 7 "Chemical Formulas and Chemical ... - Chemistry

7. What is the empirical formula for lactic acid, which has a molecular formula of $C_3H_5O_3$? a. CH_2O b. CH_2O_3 c. CH_3O_3 d. CH_3O
8. $NaCl$ is a. a molecular formula only. b. an empirical formula only. c. both an empirical formula and a chemical formula. d. both a molecular formula and an empirical formula.
- 9.

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Its molar mass is 128.18 g/mol and it contains 93.75% carbon and 6.25% hydrogen. Determine the molecular formula of naphthalene from this information. Name Class Date Name Class Date ...
CHAPTER 7 REVIEW ...

CHAPTER 7 REVIEW

1 Core Instruction Instruction and Intervention Support Chemical Formulas and Chemical Compounds Chapter Resources The

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Teacher's Edition wrap has extensive teaching support for every lesson, including Misconception Alerts, Teach from Visuals, Demonstrations, Teaching Tips, Differentiated Instruction, and Problem Solving support.

1 Core Instruction

Chapter 7 Section 1 Chemical Names and Formulas Significance of a Chemical Formula, continued •The chemical formula for an ionic compound represents one formula unit—the simplest ratio of the compound's positive ions (cations) and its negative ions (anions). •example: aluminum sulfate —Al₂(SO₄)₃

Chemical Formulas and Chemical Chapter 7 Compounds

CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 1 SHORT ANSWER 1. In a Stock system name such as iron(III) sulfate, the Roman numeral tells us (a) (b) (c) (d) 2. Answer the following questions in the space provided.

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Modern Chemistry 61 Chemical Bonding CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Write formulas for the following compounds: _____ a. copper(II) carbonate _____ b. sodium sulfite _____ c. ammonium phosphate

CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds

View Homework Help - chapter 7 from CHEM intro at Wenatchee High School. . . . _ _ . _ . _ . _ . _ . _ | 7 Q CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds M SHORT ANSWER Answer the following

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chapter 7 - I 7 Q CHAPTER 7 REVIEW Chemical Formulas and ...

Chapter 7 - Chemical Formulas and Chemical Compounds 7-1
Chemical Names and Formulas I. Significance of a Chemical Formula A. Molecular formulas 1. Number of atoms of each element in one molecule of a compound C_2H_6 = ethane (2 carbon atoms, 6 hydrogen atoms) B. Ionic Compounds 1.