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SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. c In the modern periodic table, elements are ordered (a) according to decreasing atomic mass. (b) according to Mendeleev's original design. (c) according to increasing atomic number. (d) based on when they were discovered.

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3 and write the noble-gas notation for each of the following: a. the Group 14 element in Period 4

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SHORT ANSWER Answer the following questions in the space provided. 1.

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3 Consider the neutral atom with 53 protons and 74 neutrons to answer the following questions. 53. a. What is its atomic number? 127. b. What is its mass number? atomic number. c. Is the element's position in a modern periodic table ...

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Chapter 5: The Periodic Law . Section 5.1: The History of the Periodic Table Dmitri Mendeleev (1869) – first person to

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organize the elements in a chart
Organized about 70 elements by increasing atomic mass Left blank spaces for elements which were not discovered yet . Mendeleev's Periodic Table. Periodic Table Henry Moseley (1911) - rearranged Mendeleev's periodic table to the current ...

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SECTION 5-3 SHORT ANSWER Answer the following questions in the space provided. 1. When an electron is added to a neutral atom, energy is . (a) always absorbed (c) either absorbed or released (b) always released (d) burned away 2.

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3 The energy required to remove an electron from an atom is the atom's .

Modern Chemistry Chapter 5 Review The Periodic Law Section ...

The periodic law states that the properties of elements are periodic functions of their atomic numbers. This means that the ____ determines the

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3 position of each element in the periodic table. This means that the _____ determines the position of each element in the periodic table.

Chemistry Chapter 5 Exam

The left side of the periodic table is composed of metals, which have few valence electrons. Moving to the right,

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there are more electrons and their mobility is less, reducing their metallic properties. When most of the orbitals are filled, the atoms take on nonmetallic characteristics. When orbitals are full, a noble gas configuration is attained. Metals become positive ions. When sodium ...

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Answers to Topic 5 Periodic Table from Review Book

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SECTION 1 1. _____ In the modern periodic table, elements are ordered (a) according to decreasing atomic mass. (b) according to Mendeleev's original design.

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3 Chapter: The Periodic Law Use the periodic table below to answer the questions in this Chapter Test. In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. _____ 1. Mendeleev organized the chemical ...

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Assessment Chapter Test A

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