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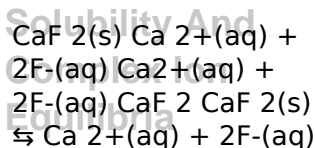
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A. Dynamic Equilibrium

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1. Equilibrium occurs when the solution is saturated B.  $K_{sp}$  (Solubility Product Constant, Solubility Product)  $K_{sp} = [\text{Ca}$

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Solubility Equilibria and  
the Solubility Product.

For a partly soluble or  
insoluble solid such as  
AgCl,  $\text{AgCl}(s) \leftrightarrow \text{Ag}^+ +$

$\text{Cl}^-$  (aq) we

define  $K_{sp}$ , the

solubility product, as.

$K_{sp} = [\text{Ag}^+][\text{Cl}^-]$

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Equilibria

Chapter 16: Solubility  
and Complex Ion

Equilibria 1. The molar  
solubility of  $PbI_2$  is  
 $1.52 \times 10^{-3} M$ .

Calculate the value of  
 $K_{sp}$  for  $PbI_2$ . A)  $3.51$   
 $10^9$  B)  $4.62 \times 10^6$  C)

$1.40 \times 10^8$  D)  $1.52 \times 10^3$   
E) none of these ANS:

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product constant MSC:  
Quantitative 2.

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### **Complex Ion**

### **Equilibria 1 ...**

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Equilibria and the

Solubility Product.

Solubility Equilibria.

Solubility product ( $K_{sp}$ )

equilibrium. constant;

has only one value for

a given.

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Solubility Equilibria.

16.1 Solubility

Equilibria and the

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Constants. When a

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slightly soluble salt such as silver chloride,  $\text{AgCl}$ , is dissolved in water, a saturated is quickly obtained, because only a very small amount of the solid dissolves, while the rest remains undissolved.

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Equilibria (2) سردم ل ا  
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### Solubility And Complex Ion Equilibria

this lecture, we continue our discussion to cover the ion product ( $Q$ ), and determining ions concentrations in the presence of many substances. We also discuss selective precipitation.

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Section 16.3 Equilibria  
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Complex Ions and Solubility Two strategies for dissolving a water-insoluble ionic solid. If the anion of the solid is a good base, the solubility is greatly increased by acidifying the solution. In cases where the anion is not sufficiently basic, the

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solubility of  $PbI_2$  is  
 $1.52 \times 10^{-3} M$ .

Calculate the value of  
 $K_{sp}$  for  $PbI_2$ . A)  $3.51 \times 10^{-9}$   
B)  $4.62 \times 10^{-6}$   
C)  $1.40 \times 10^{-8}$   
D)  $1.52 \times 10^{-3}$   
E) none  
of these 2.

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### Solubility And Equilibria 16.1

Solubility Equilibria and  
the Solubility Product

When a solid salt is added to water, it sets up an equilibrium state:

$$\text{CaF}_2(\text{s}) \rightleftharpoons \text{Ca}^{2+}(\text{aq}) + 2\text{F}^{-}(\text{aq})$$

$K_{\text{sp}} = [\text{Ca}^{2+}][\text{F}^{-}]^2$   $K_{\text{sp}}$  is the solubility product constant the amount of excess solid does not ...

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COMPLEX ION

EQUILIBRIA 669 14.

The obvious choice is that the metal ion reacts with  $\text{PO}_4^{3-}$ —and forms an insoluble phosphate salt.

**Mrs. Beierman**

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