

Chapter 14 Chemical Equilibrium

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Chapter 14 Chemical Equilibrium

Chapter 14. CHEMICAL EQUILIBRIUM 14.1 THE CONCEPT OF EQUILIBRIUM AND THE EQUILIBRIUM CONSTANT Many chemical reactions do not go to completion but instead attain a state of chemical equilibrium. Chemical equilibrium: A state in which the rates of the forward and reverse reactions are equal

Chapter 14. CHEMICAL EQUILIBRIUM

Chapter 14: Chemical Equilibrium Q1. A reaction with an equilibrium constant $K_c = 1.5 \times 10^{21}$ would consist of which of the following at equilibrium: A) approximately equal reactants and products B) some reactants and products with reactants slightly favored C) some reactants and products with products slightly favored D) essentially all reactants

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Chapter 14: Chemical Equilibrium

chemical equilibrium. a chemical reaction in which the products re-form the original.... a chemical reaction that proceeds to such an extent that at le.... a reversible reaction in which there is no longer any change i.... a state of balance in which the rate of a forward reaction equ.... reversible reaction.

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Chemical Equilibrium The state in which the rate of the forward reaction equals the rate of the reverse reaction, so that the relative concentrations of the reactants and products remain unchanged. dynamic equilibrium

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chemical equilibrium. a chemical reaction in which the products

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re-form the original.... a chemical reaction that proceeds to such an extent that at le.... a reversible reaction in which there is no longer any change i.... a state of balance in which the rate of a forward reaction equ.... reversible reaction.

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Chapter 14: Chemical Equilibrium. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. isabelagc. Key Concepts: Terms in this set (37) The rates of most chemical reactions depend on the concentrations of their reactants. Therefore, we expect the rate of the forward reaction to decrease as the concentrations of the ...

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Chapter 14: Chemical Equilibrium. [1] Which is the correct equilibrium constant expression for the following reaction? Fe. 2.

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O. 3. (s) + 3H. 2. (g) 2Fe (s) + 3H.

CHAPTER-14

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Chapter 14: Chemical Equilibrium. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. afc5xv. Terms in this set (90) reaction quotient (Q) The _____ is a measure of the progress of a reaction toward equilibrium. $Q > K$. If reaction goes to the left (towards reactants) then Q compared to K is _____

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Chapter 14: Chemical Equilibrium. Sections 14.4 - 14.7. Sarah Rodriguez. 14.4 Expressing the Equilibrium Constant in Terms of Pressure. Previously, expressed equilibrium constant w/ concentrations of reactants and products. in gaseous reactions the partial pressure of a particular gas is proportional to its

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concentration.

Chapter 14: Chemical Equilibrium

Position of Chemical Equilibrium the equilibrium position refers to the relative amounts of reactants and products in the system at the point of equilibrium a reaction with an equilibrium position that favors the products: $[\text{product}] > [\text{reactant}]$ at equilibrium equilibrium lies to the right a reaction with an equilibrium position that favors

dynamic equilibrium Chapter 14: requirements Chemical

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When the pressure is decreased by a factor of 2, the concentrations are halved, which means that the new reaction quotient is as follows: (Chapter 14.5.2) $Q = \frac{[1/2 \text{NH}_3]^2 [1/2 \text{N}_2] [1/2 \text{H}_2]^3}{1/4 [\text{NH}_3]^2 1/16 [\text{N}_2] [\text{H}_2]^3} = 4 K$.

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Two moles of gaseous products form from 2 mol of gaseous reactants.

Chapter 14.5: Factors That Affect Equilibrium - Chemistry

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Chapter 14. CHEMICAL EQUILIBRIUM 14.1 THE CONCEPT OF EQUILIBRIUM AND THE EQUILIBRIUM CONSTANT Many chemical reactions do not go to completion but instead attain a state of chemical equilibrium.

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Learning Objective. To understand what is meant by chemical equilibrium. Chemical equilibrium is a dynamic process that consists of a forward reaction, in which reactants are converted to products, and a reverse reaction, in which products are converted to reactants. At equilibrium, the forward and reverse reactions proceed at equal rates.

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Chapter 14.1: The Concept of Chemical Equilibrium ...

Chapter 14: Chemical Equilibrium Expand/collapse global location Chapter 14.2: The Equilibrium Constant ... the system described in Equation 14.2.19 will reach chemical equilibrium only if a stoichiometric amount of solid carbon or excess solid carbon has been added so that some is still present once the system has reached equilibrium.

Chapter 14.2: The Equilibrium Constant - Chemistry LibreTexts

Calculating an Equilibrium Constant from Equilibrium Concentrations. We saw in the exercise in Example 6 in Section 14.2 that the equilibrium constant for the decomposition of $\text{CaCO}_3(\text{s})$ to $\text{CaO}(\text{s})$ and $\text{CO}_2(\text{g})$ is $K = [\text{CO}_2]$. At 800°C , the concentration of CO_2 in equilibrium with solid CaCO_3 and CaO is $2.5 \times 10^{-3} \text{ M}$. Thus K at 800°C is 2.5×10^{-3} . (Remember

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that equilibrium constants ...

Chapter 14.3: Solving Equilibrium Problems - Chemistry

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